corrigé type Master I. C.M

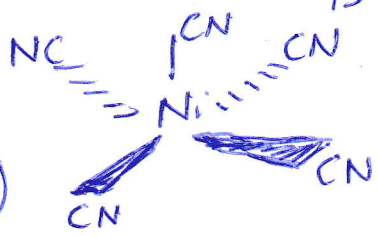
Module: chimie organométallique

Exercice 1 Décompte électronique des complexes.

a) 

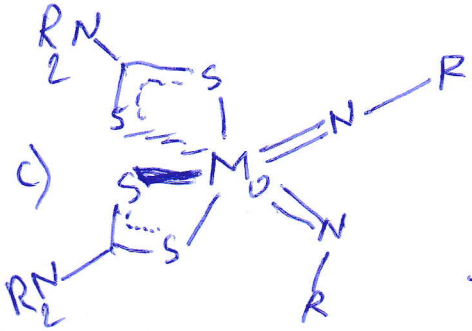
Mo 6é
allyl 3é
NO lin 3é
CO 2é
Cp 5é
+9 +1
NEV = 18é (01)

DO = +II (01)

b) 

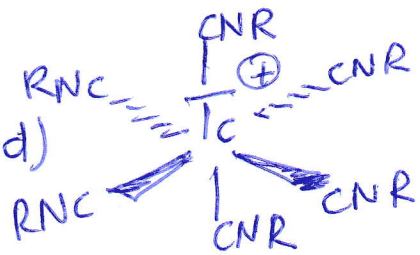
Ni 10é
5CN 5é
-9 +3
NEV = 18é (01)

DO = +II (01)

c) 

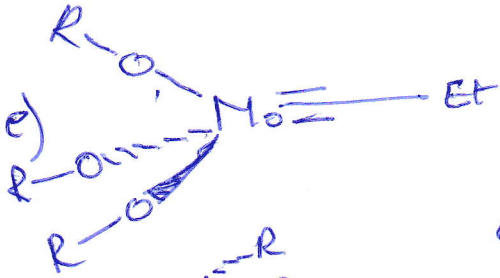
Mo 6é
2 $\eta^2 S_2$ 6é
RN coudé 2
RN lin 4
NEV = 18é (01)

DO = +VI (01)

d) 

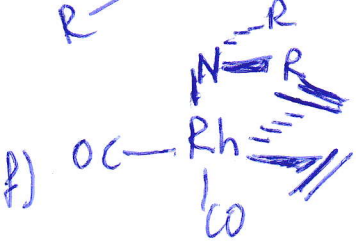
Tc 7é
6CNR 12é
-9 +1
NEV = 18é (01)

DO = +I (01)

e) 

Mo 6é
3RO (lin) 9é
Et + C 3é
NEV = 18é (01)

DO = +VI (01)

f) 

Rh 9é
2CO 4é
2(C_2H_4) 4é
NR_2 1é
NEV = 18é (1)

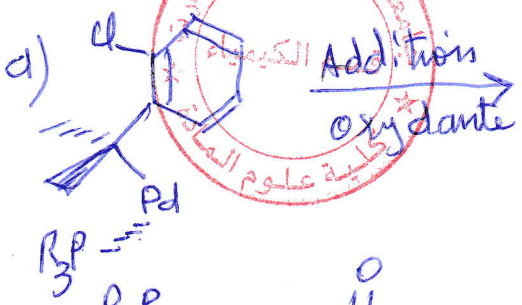
DO = +I (1)

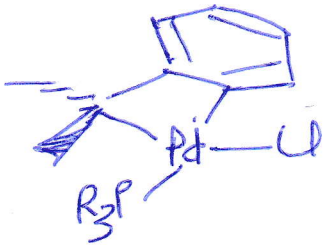
g) $(\eta^3\text{-}CH_2CHCHMe)Fe(CO)_3$ (1)

Fe 8é
Allyl 3é
3CO 6é
Ø 1é
NEV = 18é (01)

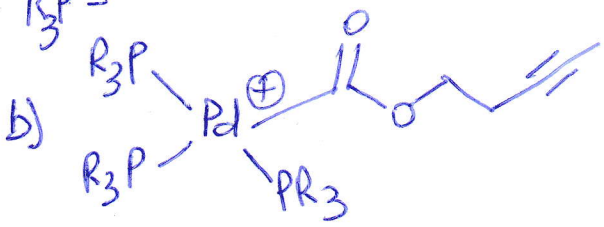
DO = +II (01)

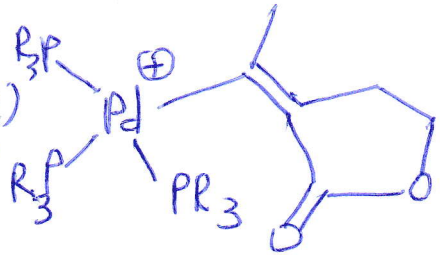
Exercice 2 :

a) 

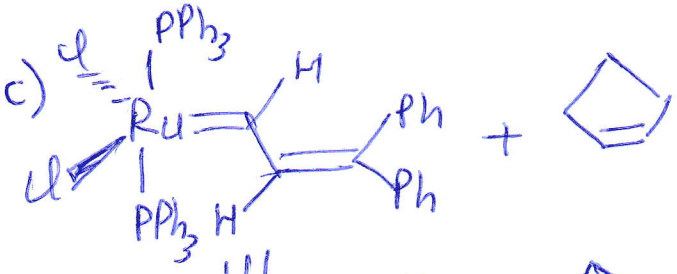


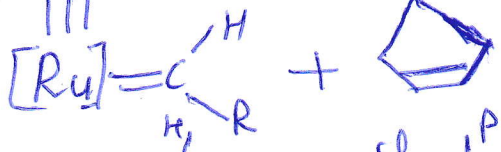
(1)

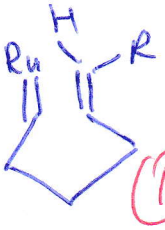
b) 



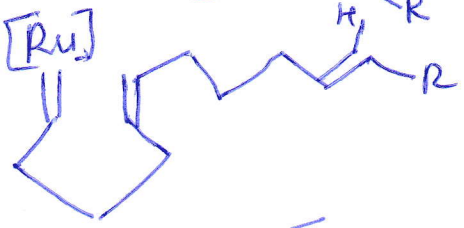
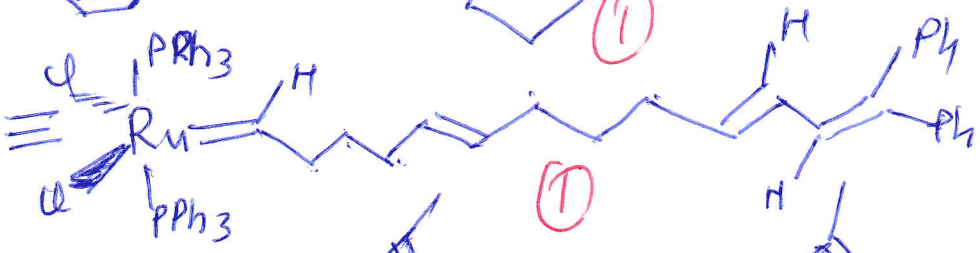
(1)

c) 

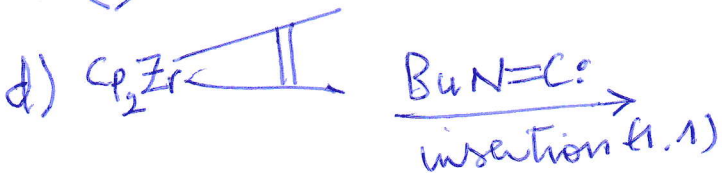




(1)

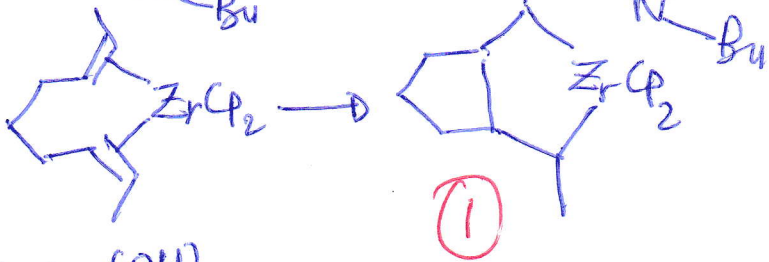



(1)

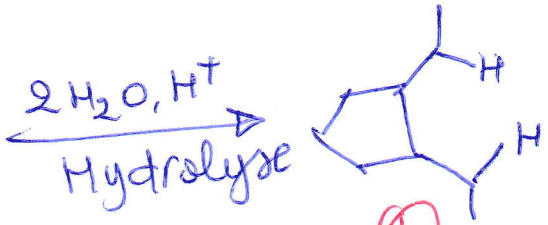
d) 

(1)

e) 



(1)



(1)

Exercice 3 Mécanisme